

Chemical Reactions and Balancing Chemical Equations

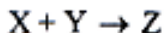
Created by faculty and staff . South Campus Learning Center, Science Lab
01-06-06 Permission to copy and use is granted to all FCCJ staff
provided this copyright label is displayed.

For more information, visit the Learning Services web site: www.fccj.edu/lc

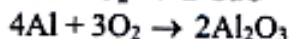
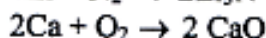
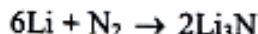
Types of Chemical Reactions

Web Link: <http://web.fccj.edu/~ksanchez/1032/wksheet/Reactions.htm>

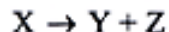
COMBINATION - Two or more reactants form a single product.



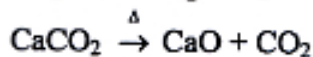
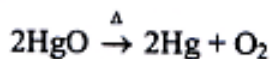
Examples:



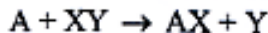
DECOMPOSITION - Single reactant forms two or more products.



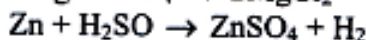
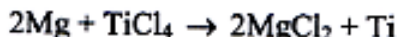
Example:



SINGLE REPLACEMENT - One substance replaces another in a compound.



Example:

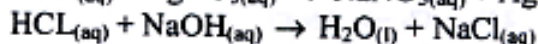
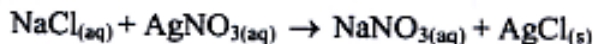


(Note: Check "Activity Series" to see if some replacements will occur.)

DOUBLE REPLACEMENT - Two compounds "change partners".

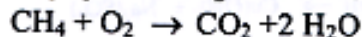
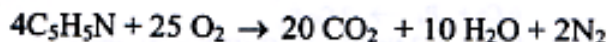


Example:



COMBUSTION - Generally involves a "fuel" being "burned" in oxygen and production of CO_2 and H_2O .

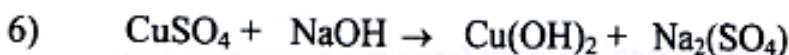
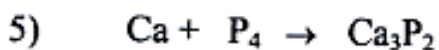
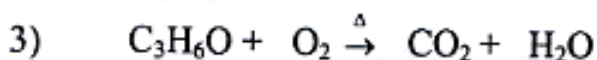
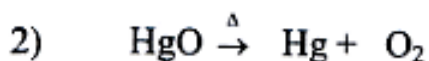
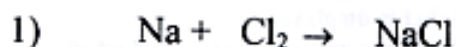
Example:



Balancing Chemical Equations

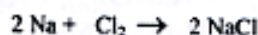
Web Link: <http://www.wfu.edu/~ylwong/balanceeq/balanceeq.html>

Using the examples from the websites above, name and balance the following reactions.

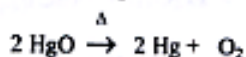


Answers:

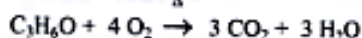
1) Combination,



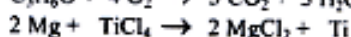
2) Decomposition,



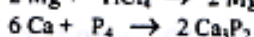
3) Combustion,



4) Single Replacement,



5) Combination,



6) Double Replacement,

